



## **Ecofriendly extract of *Citrus reticulata* as corrosion inhibitor for Copper in NaOH Solution**

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**Abstract:** The effect of *Citrus reticulata* extracts on the corrosion of copper in 1M NaOH solution was studied using weight loss method at various temperatures varying from 298 to 328 K. Experimental data revealed that *Citrus reticulata* extract acted as an inhibitor in the alkaline environment. It was found that the inhibition efficiency increased with an increase in *Citrus reticulata* extract concentration. Adsorption enthalpies were determined and discussed. Effect of temperature was also investigated and activation parameters were evaluated.

**Keywords:** Citrus reticulate, Inhibition efficiency, Copper, Weight loss.

### **1. Introduction**

Copper has been one of the preferred metals in industry owing to its excellent electrical and thermal conductivities, good mechanical workability, and its relatively noble properties<sup>1</sup>. In spite of the fact that copper is a relatively noble metal, it reacts easily in oxygen containing electrolytes. However, wide industrial application of copper has been based on its corrosion stability, which is the result of formation of an oxide/hydroxide layer on the metal surface. The oxide layer has, in general, a duplex structure made up of an inner Cu<sub>2</sub>O layer followed by CuO and then a Cu(OH)<sub>2</sub> layer depending on the electrode potential<sup>1</sup>. Thin passivity layer formed on copper in neutral and alkaline solution has attracted considerable interest particularly in the corrosion catalysis and double layer structure research<sup>2</sup>.

Some of these corrosion inhibitors are however, toxic to the environment. This has prompted the search for green corrosion inhibitors that are non-toxic and eco-friendly for metals and alloys in acidic and alkaline solutions<sup>3</sup>. These green corrosion inhibitors have been found to have centre for  $\pi$ -electrons and functional groups (such as  $-C=C-$ ,  $-OR$ ,  $-OH$ ,  $-NR_2$ ,  $-NH_2$  and  $-SR$ ) which provide electrons that facilitate the adsorption of the inhibitor on the metal surface.

Research on the use of plant extracts as corrosion inhibitors for metals/alloys in acid or alkaline media has therefore been intensified<sup>3-10</sup>. This is because plants are rich sources of naturally occurring chemical compounds that are environmentally acceptable, cheap and readily available.

The aim of the present work is to study the inhibition of copper surface by addition of *Citrus reticulata* extract as inhibitor in NaOH solution by using weight loss technique.

### **2. Materials and methods**

The weight loss measurements were carried out in a test tube placed in a thermostat water bath. The solution volume was 10 mL. The used copper coupons had a rectangular form (length = 1 cm, width = 1 cm, thickness = 0.03 cm). Prior to all measurements, the coupons were first polished successively with metallographic emery paper of increasing fineness up to 1200 grits. The coupons were then washed with doubly

distilled water, degreased with acetone, washed using doubly distilled water again and finally, dried with paper tissue at room temperature.

## 2.1. Inhibitor material

A Stock solution of the inhibitor material was prepared by refluxing 12.5 g of dry *Citrus reticulata* powder with 250 mL of 1M NaOH for 3 hours. The refluxed solution was allowed to stand overnight and filtered through ordinary filter paper. From this solution, different concentrations of inhibitor solutions ranging from 0.1% to 0.7% were diluted.

## 2.2. Weight loss method

Pre-weighed copper specimens (in triplicate) were suspended for 1 hour in 1M NaOH with and without the inhibitor in different volume ranging from 1 to 7 mL of extract. After the specified time the coupons were removed from test solution, thoroughly washed with acetone solution and de-ionised water, dried well and then reweighed.

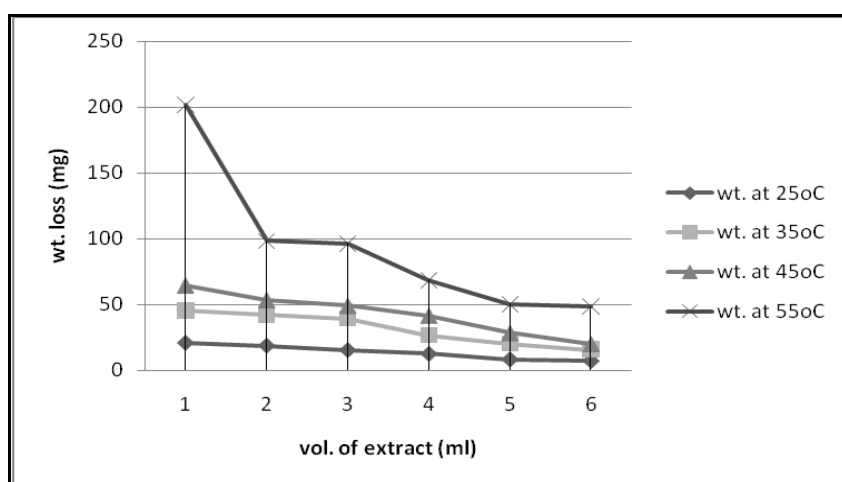
## 3. Results and discussion

The plot of weight loss (mg) of copper coupon versus inhibitor volume for 60 min immersion period at different temperatures is shown in **Fig. 1**. Weight loss as a function of inhibitor volumes decreased gradually. This indicates that *Citrus reticulata* extract inhibits the corrosion of copper in 1 M NaOH solution. The values of percentage inhibitor efficiency (%I) for various concentrations of the inhibitor was determined for 1 hour immersion periods by using the following equation:

$$\%I = [W_u - W_b/W_u] \times 100$$

where  $W_u$  and  $W_b$  are the uninhibited and inhibited weight losses, respectively. Assuming a direct relationship between inhibition efficiency (%I) and surface coverage ( $\theta$ ) for different inhibitor concentrations, the degree of surface coverage ( $\theta$ ) was calculated by using the following relationship:

$$\theta = \%I/100$$



**Figure 1.** The plot of weight loss of copper coupon versus volume of *Citrus reticulata* extract in 1M NaOH for 60 min. immersion period at different temperatures.

**Table 1** collects the percentage inhibition efficiency for 60 min immersion periods at 25, 35, 45 and 55 °C. The percentage inhibition efficiency values, as presented in Table 1 for triplicate copper specimens. As seen in **Table 1**, the percentage inhibition efficiency values increase with increasing extract concentration, but decrease with increasing temperature. The highest inhibition efficiency of 86.49% was obtained at 7 mL of extract at 25 °C. This result suggests that increase in extract concentration increases the number of inhibitor molecules adsorbed onto copper surface and reduces the surface area that is available for the direct base attack

on the metal surface. The inhibitive effect of *Citrus reticulata* is ascribed to the presence of organic compounds in the extract. *Citrus reticulata* is rich in several organic compounds of high molecular weight with heteroatom and  $\pi$  centers in their molecular structures.

**Table 1.** Inhibition efficiency of *Citrus reticulata* extract on copper in 1M NaOH for 60 min immersion period at different temperatures

Extract	%I at 25 °C	%I at 35 °C	%I at 45 °C	%I at 55 °C
1	54.03	46.20	33.26	18.91
2	63.45	58.14	46.74	28.87
3	73.80	71.45	56.65	41.65
5	83.17	79.08	74.03	60.54
7	85.49	82.34	78.13	72.74

The inhibition effect of *Citrus reticulata* may be due to the presence of these organic compounds in the extract. Since *reticulata* contains several compounds, synergistic and antagonistic effects may play an important role on the inhibition efficiency of *Citrus reticulata* as an inhibitor. Organic compounds having centers for  $\pi$  electrons and functional groups of O have been reported as corrosion inhibitors for copper in basic solutions<sup>12</sup>. The adsorption of these compounds on copper surface reduces the surface area that is available for the attack of the aggressive ion from the basic solution. As seen in **Fig. 1** the weight losses decrease with increase in extract concentration due to higher degree of surface coverage,  $\theta$  as a result of enhanced inhibitor adsorption. Similar view has been reported previously<sup>5-8</sup>. Also, Fig. 2 confirms that the inhibition is due to the adsorption of the active organic compounds onto metal surface. This is because a straight line is obtained when  $C/\theta$  is plotted against  $C$  and the linear correlation coefficient of the fitted data is close to 1, indicating that the adsorption of the inhibitor molecules obey the Langmuir's adsorption isotherm expressed as<sup>9</sup>.

$$C/\theta = C + 1/K$$

Where  $C$  is the inhibitor concentration and  $K$  the equilibrium constant for the adsorption/desorption process of the inhibitor molecules on the metal surface.

The effect of increase in solution temperature from 25 to 55 °C on the inhibitor efficiency is summarized in table increased in temperature with produce a decrease in the inhibitor efficiency and this suggest that the process of adsorption of the inhibitor molecules is physical in nature. The apparent activation energy,  $E_a$  of the corrosion reaction was calculated by using the Arrhenius equation.

$$\text{Log}(R_{\text{Corr1}}/R_{\text{Corr2}}) = E_a/2.303R(1/T_1-1/T_2)$$

Where  $R_{\text{Corr1}}$  and  $R_{\text{Corr2}}$  are corrosion rates at temperature  $T_1$  and  $T_2$ , respectively. The corrosion rate ( $R_{\text{Corr}}$ ,  $\text{mg cm}^{-2} \text{ h}^{-1}$ ) over the exposure time were calculated as follows<sup>13-15</sup>.

$$\Delta W = (W_1 - W_2)/AR_{\text{Corr}} = \Delta W/t$$

Where,  $W_1$  is the original weight (mg) of the copper specimen,  $W_2$  the weight after immersion in the test electrolyte,  $A$  is the area in  $\text{cm}^2$ , and  $t$  is the time of exposure in h.

The heat of adsorption,  $Q_{\text{ads}}$  was calculated by using the following equation<sup>15</sup>.

$$Q_{\text{ads}} = 2.303 R[\text{log}(\theta_2/1-\theta_2) - \text{log}(\theta_1/1-\theta_1)] \times (T_2T_1/T_2-T_1)$$

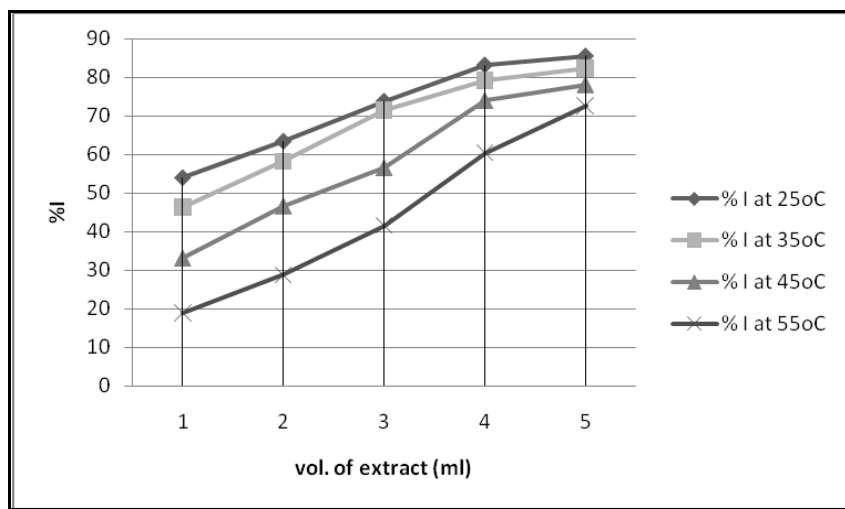
Where  $\theta_1$  and  $\theta_2$  are the values of surface coverage at temperatures  $T_1$  and  $T_2$ .

Arrhenius plot of logarithmic of corrosion rate against the reciprocal of absolute temperature ( $1/T$ ) can be shown graphically. The values of activation energy ( $E_a$ ) and heat of adsorption ( $Q_{\text{ads}}$ ) are presented in **Table 2**. The observed increase in activation energy in the presence of inhibitor from 53.06 to 71.97  $\text{kJ mol}^{-1}$ , with attendant decrease in inhibition efficiency of the inhibitor as temperature increases, suggests physical adsorption of the inhibitor molecules on the copper surface. This is in accordance with the findings of other workers<sup>11,15</sup>. The negative values of the  $Q_{\text{ads}}$  of adsorption and the high values of the adsorption constant indicate a spontaneous adsorption of these inhibitors on copper. This means that the inhibitive action of these substances results from the physical adsorption of these molecules on the surface of copper.

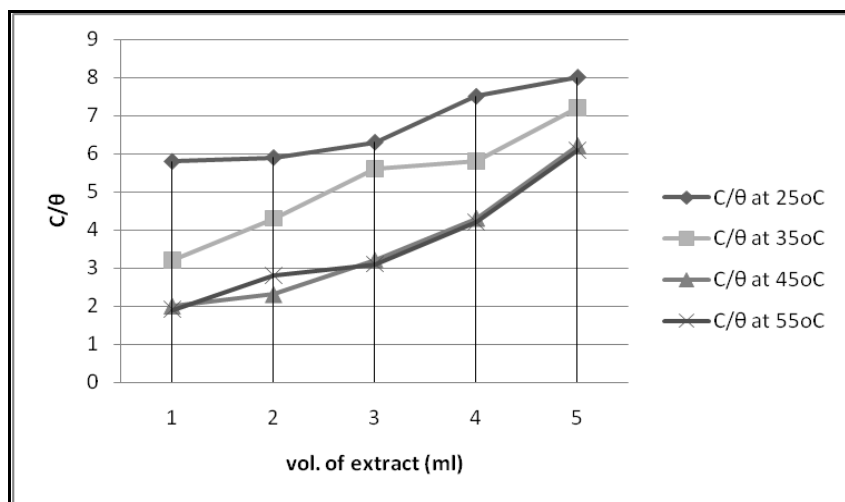
**Table 2.** The activation energy ( $E_a$ ) kJ/mol and  $Q_{ads}$  (kJ/mol) of the copper with and without inhibitions

(% v/v)	$E_a$ (kJ/mol)	$Q_{ads}$ (kJ/mol)
0	53.06	-
1	58.30	-37.98
2	61.19	-29.57
3	63.10	-32.35
5	63.47	-9.13
7	71.97	-22.54

**Figure 2.** Inhibition efficiency of extract on copper in 1M NaOH for 60 min immersion period at different temperatures.



**Figure 3.** Langmuir isotherm for the adsorption on the copper surface of *Citrus reticulata* extract at different temperature.



**Conclusion**

The extract of *Citrus reticulata* inhibits the corrosion of copper in 1 M NaOH solutions, with inhibition efficiency of 86.49% at 7 mL of extract and the % inhibition efficiency decreased with increase in temperature.

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## References

1. Wilhelm S.M., Tanizawa Y., Liu, Chang-Yi and Hackerman N., A photoelectrochemical investigation of semiconducting oxide films on copper, *Corrosion Science*, 22(8), 791 (1982).
2. Kautek W. and Gordon J.G., XPS studies of anodic surface films on copper electrodes, *Journal of the Electrochemical Society*, 137(9), 2672 (1990).
3. Oguzie E.E., Corrosion inhibition of aluminum in acidic and alkaline media by *Sansevieria trifasciata* extract, *Corrosion Science*, 49(3), 1527 (2007).
4. Abdel-Gaber A.M., Khamis E. and Abo-ElDahab H., Inhibition of aluminium corrosion in alkaline solutions using natural compound, *Materials Chemistry and Physics*, 109(2-3), 297 (2008).
5. Tripathi R., Chaturvedi A. and Upadhyay R.K., Corrosion Inhibitory Effects of Some Substituted Thiourea on Mild Steel in Acid Media, *Res. J. Chem. Sci.*, 2(2), 18-27 (2012).
6. Abiola Olusegun K., Otaigbe J.O.E. and Kio O.J., *Gossypium hirsutum* L. extracts as green corrosion inhibitor for aluminum in NaOH solution, *Corrosion Science*, 51(8), 1879 (2009).
7. Manivannan M. and Rajendran S., Corrosion Inhibition of Carbon steel by Succinic acid – Zn (II) system, *Res. J. Chem. Sci.*, 1(8), 1-7 (2011).
8. Onen A.I., Maitera O.N., Joseph J. and Ebenso E.E., Corrosion inhibition potential and adsorption behavior of Bromophenol blue and Thymol blue dyes on mild steel in acidic medium, *International Journal of Electrochemical Science*, 6(7), 2884 (2011).
9. Shylesha B.S., Venkatesha T.V. and Praveen B.M., Corrosion Inhibition Studies of Mild Steel by New Inhibitor in Different Corrosive Medium, *Res. J. Chem. Sci.*, 1(7), 46-50 (2011).
10. Nutan Kumpawat, Alok Chaturvedi and Upadhyay R.K., Corrosion Inhibition of Mild Steel by Alkaloid Extract of *Ocimum Sanctum* in HCl and HNO<sub>3</sub> Solution, *Res.J.Chem.Sci.*, 2(5), 51-56, (2012).
11. Ortal D., Hugo E.G., Shlomo G. and Margalit B., Antioxidant activity of 1,3-dicaffeoylquinic acid isolated from *Inula viscosa*, *Food Research International*, 42, 1273 (2009).
12. Desai M.N., Corrosion inhibitors for copper, *Materials and Corrosion*, 23 (6), 483 (1972).
13. James A.O. and Alaranta O., Inhibition of Corrosion of Zinc in Hydrochloric acid solution by Red Onion Skin Acetone Extract, *Res. J. Chem. Sci.*, 1(1), 31-37 (2011).
14. Sherif E.M., Effects of 2-amino-5-(ethylthio)-1,3,4-thiadiazole on copper corrosion as a corrosion inhibitor in 3% NaCl solutions, *Applied Surface Science*, 252, 8615 (2006).
15. Oguzie E.E., Corrosion inhibition of aluminium in acidic and alkaline media by *Sansevieria trifasciata* extract, *Corrosion Science*, 49, 1527 (2007).

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